

CHM 130LL: Formula Writing and Naming Key

EXERCISE

Write the formula for each of the following compounds. Consult your list of polyatomic ions when necessary.

1. potassium hydroxide KOH	13. lithium cyanide LiCN
2. sodium carbonate Na ₂ CO ₃	14. barium hydroxide Ba(OH) ₂
3. ammonium bromide NH ₄ Br	15. phosphorus trichloride PCl ₃
4. silver chromate Ag ₂ CrO ₄	16. copper (II) bromide CuBr ₂
5. magnesium nitride Mg ₃ N ₂	17. sodium acetate NaC ₂ H ₃ O ₂
6. carbon tetrachloride CCl ₄	18. sulfur hexafluoride SF ₆
7. dinitrogen oxide N ₂ O	19. diphosphorus pentasulfide P ₂ S ₅
8. nitric acid HNO ₃ (aq)	20. mercury (II) phosphate Hg ₃ (PO ₄) ₂
9. lead (II) chloride PbCl ₂	21. sulfuric acid H ₂ SO ₄ (aq)
10. cobalt (III) sulfate Co ₂ (SO ₄) ₃	22. calcium sulfite CaSO ₃
11. iron (II) phosphide Fe ₃ P ₂	23. zinc chromate ZnCrO ₄
12. ammonium sulfate (NH ₄) ₂ SO ₄	24. hydrochloric acid HCl(aq)

EXERCISE

Write the name for each of the following compounds. Consult your list of polyatomic ions when necessary.

1. Na_2SO_4 sodium sulfate	13. SO_3 sulfur trioxide
2. CF_4 carbon tetrafluoride	14. Ag_2S silver sulfide
3. AlPO_4 aluminum phosphate	15. KCN potassium cyanide
4. CaI_2 calcium iodide	16. NH_4I ammonium iodide
5. MgCrO_4 magnesium chromate	17. $\text{Hg}(\text{NO}_3)_2$ mercury(II) nitrate
6. FeBr_3 iron(III) bromide	18. $\text{Al}(\text{NO}_2)_3$ aluminum nitrite
7. $\text{Co}(\text{OH})_2$ cobalt (II) hydroxide	19. FeP iron(III) phosphide
8. $\text{HBr}(\text{aq})$ hydrobromic acid	20. $\text{Mg}(\text{MnO}_4)_2$ magnesium permanganate
9. ZnSO_3 zinc sulfite	21. CuSO_4 copper(II) sulfate
10. $\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$ barium acetate	22. Ni_3N_2 nickel (II) nitride
11. LiNO_3 lithium nitrate	23. P_2O_5 diphosphorus pentoxide
12. $\text{Cr}_2(\text{CO}_3)_3$ chromium(III) carbonate	24. $\text{H}_3\text{PO}_4(\text{aq})$ phosphoric acid