

CHM 130LL:
Formula Writing and Naming

Name: _____

Partner: _____

Section Number: _____

Write the formula for each of the following compounds. Consult your list of polyatomic ions when necessary.

1. potassium hydroxide	13. lithium cyanide
2. sodium carbonate	14. barium hydroxide
3. ammonium bromide	15. phosphorus trichloride
4. silver chromate	16. copper (II) bromide
5. magnesium nitride	17. sodium acetate
6. carbon tetrachloride	18. sulfur hexafluoride
7. dinitrogen oxide	19. diphosphorus pentasulfide
8. nitric acid	20. mercury (II) phosphate
9. lead (II) chloride	21. sulfuric acid
10. cobalt (III) sulfate	22. calcium sulfite
11. iron (II) phosphide	23. zinc chromate
12. ammonium sulfate	24. hydrochloric acid

EXERCISE

EXERCISE

Write the name for each of the following compounds. Consult your list of polyatomic ions when necessary.

1. Na_2SO_4	13. SO_3
2. CF_4	14. Ag_2S
3. AlPO_4	15. KCN
4. CaI_2	16. NH_4I
5. MgCrO_4	17. $\text{Hg}(\text{NO}_3)_2$
6. FeBr_3	18. $\text{Al}(\text{NO}_2)_3$
7. $\text{Co}(\text{OH})_2$	19. FeP
8. $\text{HBr}(\text{aq})$	20. $\text{Mg}(\text{MnO}_4)_2$
9. ZnSO_3	21. CuSO_4
10. $\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$	22. Ni_3N_2
11. LiNO_3	23. P_2O_5
12. $\text{Cr}_2(\text{CO}_3)_3$	24. $\text{H}_3\text{PO}_4(\text{aq})$

Note: A key for these exercises will be available in the CHM 130 cabinet next week, so you can check your answers.