

## Solubility Practice Problems

1. Write the reaction and  $K_{sp}$  expression for barium sulfate.
2. The  $K_{sp}$  for calcium fluoride is  $3.9 \times 10^{-11}$ . Calculate its solubility. Write the reaction for full credit.
3. Calculate the molar solubility of calcium fluoride in a solution containing 0.010M calcium nitrate. Write the reaction for full credit.
4. Will a precipitate form if 0.10 L of  $8.0 \times 10^{-3}$ M lead(II) nitrate is added to 0.40 L of  $5.0 \times 10^{-3}$ M sodium sulfate? Write all balanced reactions for full credit.