$$°F = (1.8 \times °C) + 32 \qquad °C = \frac{(°F - 32)}{1.8} \qquad K = °C + 273$$

1 inch = 2.54 cm   1 lb = 454 g   1 qt = 946 mL

$$M = \frac{moles\ solute}{L\ solution} \qquad mass\ \%\ concentration = \frac{mass\ of\ solute}{mass\ of\ solution} \times 100\%$$
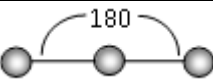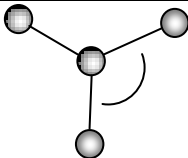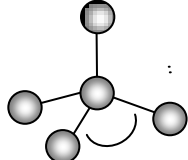
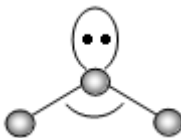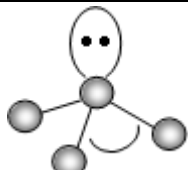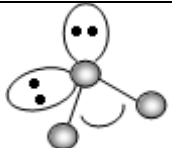1 mole = 6.02×10^{23} particles   1 mole of gas = 22.4 L at STP

1 atm = 760 torr = 760 mm Hg = 14.7 psi $\dfrac{P_1 V_1}{T_1} = \dfrac{P_2 V_2}{T_2}$

CHM130 Table of Molecular Shapes and Bond Angles

General formula	MOLECULAR GEOMETRY	NAME of SHAPE	Bond Angles
AB$_2$		linear	180˚
AB$_3$		trigonal planar	120˚
AB$_4$		tetrahedral	109.5˚
AB$_2$E		bent	<120˚
AB$_3$E		trigonal pyramid	<109.5˚
AB$_2$E$_2$		bent	<109.5˚