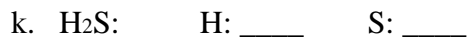
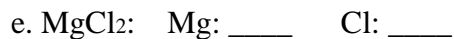
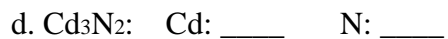
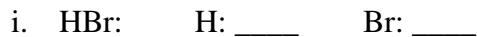
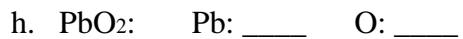
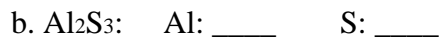
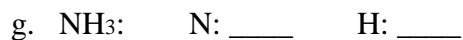
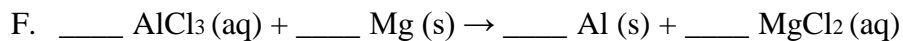
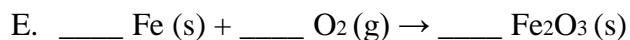
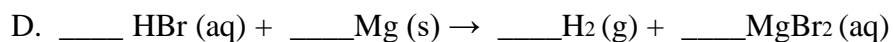
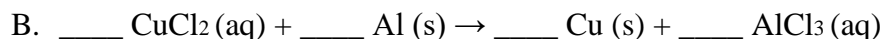
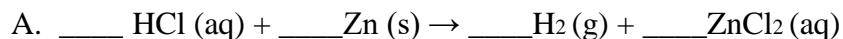


CHM 130: Redox Practice Problems

1. Determine the charge for each of the following:



2. For each of the following, Balance the equation, Identify the reactant oxidized and the reactant reduced.

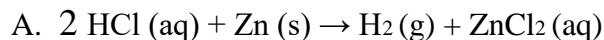


CHM 130: Redox Practice Problems Key

1. Determine the charge for each of the following:

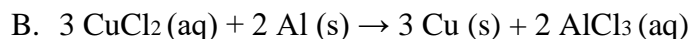
- | | |
|--|--|
| a. O ₂ : O: 0 | g. NH ₃ : N: -3 H: +1 |
| b. Al ₂ S ₃ : Al: +3 S: -2 | h. PbO ₂ : Pb: +4 O: -2 |
| c. P ₄ : P: 0 | i. HBr: H: +1 Br: -1 |
| d. Cd ₃ N ₂ : Cd: +2 N: -3 | j. K: K: 0 |
| e. MgCl ₂ : Mg: +2 Cl: -1 | k. H ₂ S: H: +1 S: -2 |
| f. Na: Na: 0 | l. I ₂ : I: 0 |

2. For each of the following, Balance the equation, Identify the reactant oxidized and the reactant reduced.



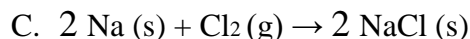
H⁺ in HCl(aq) is the reactant reduced.

Zn(s) is the reactant oxidized.



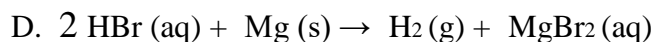
Cu²⁺ in CuCl₂ (aq) is the reactant reduced.

Al(s) is the reactant oxidized.



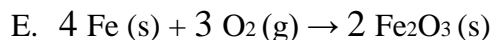
Cl₂(g) is the reactant reduced.

Na(s) is the reactant oxidized.



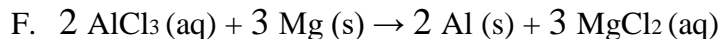
H⁺ in HBr(aq) is the reactant reduced.

Mg(s) is the reactant oxidized.



O₂(g) is the reactant reduced.

Fe(s) is the reactant oxidized.



Al³⁺ in AlCl₃(aq) is the reactant reduced.

Mg(s) is the reactant oxidized.